

**Natural gas** consists of mostly methane and is associated with petroleum deposits

**Natural law** a statement that expresses generally observed behavior

**Net ionic equation** an equation for a reaction in solution, representing strong electrolytes as ions and showing only those components that are directly involved in the chemical change

**Network solid** an atomic solid containing strong directional covalent bonds

**Neutralization reaction** an acid – base reaction

**Neutron** a particle in the atomic nucleus with a mass approximately equal to that of the proton but with no charge

**Noble gas** a Group 8 element

**Nonelectrolyte** a substance that, when dissolved in water, gives a nonconducting solution

**Nonmetal** an element that does not exhibit metallic characteristics. Chemically, a typical nonmetal accepts electrons from a metal

**Normal boiling point** the temperature at which the vapor pressure of a liquid is exactly one atmosphere; the boiling temperature under one atmosphere of pressure

**Normal melting/freezing point** the melting/freezing point of a solid at a total pressure of one atmosphere

**Normality** the number of equivalents of a substance dissolved in a liter of solution

**Nuclear atom** the modern concept of the atom as having a dense center of positive charge (the nucleus) and electrons moving around the nucleus

**Nuclear transformation** the change of one element into another

**Nucleon** a particle in an atomic nucleus, either a neutron or a proton

**Nucleus** the small, dense center of positive charge in an atom

**Nuclide** the general term applied to each unique atom; represented by X, where X is the symbol for a particular element

