

Name: _____
Hour: _____ Date: _____

Chemistry: *Ionic Formulas (Binary, Polyatomic, Transition Metals)*

Write the formula for each of the following compounds.

- | | |
|---------------------------|-----------|
| 1. sodium hydroxide | 1. _____ |
| 2. mercury (II) sulfate | 2. _____ |
| 3. lead (II) phosphate | 3. _____ |
| 4. ammonium sulfide | 4. _____ |
| 5. aluminum chlorate | 5. _____ |
| 6. copper (I) carbonate | 6. _____ |
| 7. manganese (IV) oxide | 7. _____ |
| 8. manganese (II) sulfate | 8. _____ |
| 9. iron (III) oxide | 9. _____ |
| 10. magnesium nitrate | 10. _____ |
| 11. calcium sulfide | 11. _____ |
| 12. potassium oxide | 12. _____ |
| 13. magnesium chloride | 13. _____ |
| 14. chromium (III) oxide | 14. _____ |
| 15. gold (III) bromide | 15. _____ |
| 16. beryllium fluoride | 16. _____ |
| 17. yttrium (III) sulfide | 17. _____ |
| 18. rubidium oxide | 18. _____ |
| 19. aluminum nitrate | 19. _____ |
| 20. iron (II) phosphide | 20. _____ |
| 21. ferric nitrate | 21. _____ |
| 22. ferrous nitrate | 22. _____ |
| 23. plumbous chloride | 23. _____ |
| 24. stannic hydroxide | 24. _____ |
| 25. niobium (V) oxide | 25. _____ |