# Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Hour: \_\_\_\_ Date: \_\_\_\_\_\_\_\_\_\_\_

# Chemistry: *Stoichiometry – Problem Sheet 1*

*Directions*: *Solve each of the following problems. Show your work, including proper units, to earn full credit.*

1. Silver and nitric acid react according to the following balanced equation:

3 Ag(s) + 4 HNO3(aq) 🡪 3 AgNO3(aq) + 2 H2O(l) + NO(g)

A. How many moles of silver are needed to react with 40 moles of nitric acid?

B. From the amount of nitric acid given in Part A, how many moles of silver nitrate will be produced?

C. From the amount of nitric acid given in Part A, how many moles of water will be produced?

D. From the amount of nitric acid given in Part A, how many moles of nitrogen monoxide will be made?

2. Given the balanced equation: 2 N2H4(l) + N2O4(l) 🡪 3 N2(g) + 4 H2O(g)

A. How many moles of dinitrogen tetrahydride are required to produce 57 moles of nitrogen?

B. How many moles of dinitrogen tetroxideare required to produce 57 moles of nitrogen?

C. How many moles of waterare produced when 57 moles of nitrogen are made?

3. Calculate the mass of aluminum oxide produced when 3.75 moles of aluminum burn in oxygen.

Answers: 1A. 30 mol Ag 1C. 20 mol H2O 2A. 38 mol N2H4 2C. 76 mol H2O

1B. 30 mol AgNO3 1D. 10 mol NO 2B. 19 mol N2O4 3. 191 g Al2O3

4. At a very high temperature, manganese is isolated from its ore, manganomanganic oxide, via the following balanced equation:

3 Mn3O4(s) + 8 Al(s) 🡪 4 Al2O3(s) + 9 Mn(s)

A. How many manganese atoms are liberated if 54.8 moles of Mn3O4 react with excess aluminum.

B. How many moles of aluminum oxide are made if 3580 g of manganomanganic oxide are consumed?

C. How many moles of manganomanganic oxide will react with 5.33 x 1025 atoms of aluminum?

D. If 4.37 moles of aluminum are consumed, how many molecules of aluminum oxide are produced?

5. Camels store the fat tristearin (C57H110O6) in the hump. Besides being a source of energy, the fat is a source of water for the camel because when the fat is burned, the following reaction occurs:

2 C57H110O6(s) + 163 O2(g) 🡪 114 CO2(g) + 110 H2O(l)

A. At STP, what volume of oxygen is required to consume 0.64 moles of tristearin?

B. At STP, what volume of carbon dioxide is produced in Part A?

C. If 22.4 L of oxygen is consumed at STP, how many moles of water are produced?

D. Find the mass of tristearin required to produce 55.56 moles of water (about 1 liter of liquid water).

Answers: 4A. 9.9 x 1025 atoms Mn 4C. 33.2 mol Mn3O4 5A. 1168 L O2 5C. 0.675 mol H2O

4B. 20.9 mol Al2O3 4D. 1.3 x 1024 m’cules Al2O3 5B. 817 L CO2 5D. 899 g C57H110O6