Strong / Weak Acid vs. Concentrated and Dilute

Concentration: Molarity, Molality, Normality

Solutions Guide (#1-5)

Solutions (#1-5)

Solutions (#1-10) front and (#1-3) back

Chemistry: Molarity and Stoichiometry (#1-4)

pH scale

Acids: sour taste, turn litmus paper red

Arrhenius’s Definition (H1+ as only ion in water)

Bronsted-Lowry - a proton donor

Bases: bitter taste, turn litmus paper blue

Arrhenius’s Definition (OH1- as only negative ion)

Bronstred-Lowry - a proton acceptor

Strength of Acid Ka = acid dissociation constant

Kw = 1 x 10-14 Kw = [H1+][OH1-]

Indicators

Buffers

Common acids and bases

Acid + Base 🡪 Salt + Water Neutralization

Acids and Bases (#1-10)

Practice Problems (#1-13) Acids and Bases

# Aqueous Acids and Bases – Titration (#1-8)

LeChattelier’s Principle